

3220/202

SCOTTISH
CERTIFICATE OF
EDUCATION
1996

FRIDAY, 17 MAY
1.30 PM – 4.00 PM

PHYSICS
HIGHER GRADE
Paper II

Read carefully

- 1 All questions should be attempted.
- 2 Enter the question number clearly in the margin beside each question.
- 3 Any necessary data will be found in the Data Sheet on page two.
- 4 Care should be taken not to give an unreasonable number of significant figures in the final answers to calculations.
- 5 Square-ruled paper (if used) should be placed inside the front cover of the answer book for return to the Examination Board.



DATA SHEET
COMMON PHYSICAL QUANTITIES

Quantity	Symbol	Value	Quantity	Symbol	Value
Speed of light in vacuum	c	$3.00 \times 10^8 \text{ m s}^{-1}$	Mass of electron	m_e	$9.11 \times 10^{-31} \text{ kg}$
Charge on electron	e	$-1.60 \times 10^{-19} \text{ C}$	Mass of proton	m_p	$1.673 \times 10^{-27} \text{ kg}$
Gravitational acceleration	g	9.8 m s^{-2}			
Planck's constant	h	$6.63 \times 10^{-34} \text{ J s}$			

REFRACTIVE INDICES

The refractive indices refer to sodium light of wavelength 589 nm and to substances at a temperature of 273 K.

Substance	Refractive index	Substance	Refractive index
Diamond	2.42	Glycerol	1.47
Glass	1.51	Water	1.33
Ice	1.31	Air	1.00
Perspex	1.49		

SPECTRAL LINES

Element	Wavelength/nm	Colour	Element	Wavelength/nm	Colour
Hydrogen	656	Red	Cadmium	644	Red
	486	Blue-green		509	Green
	434	Blue-violet		480	Blue
	410	Violet	<i>Lasers</i>		
	397	Ultraviolet	<i>Element</i>	<i>Wavelength/nm</i>	<i>Colour</i>
	389	Ultraviolet	Carbon dioxide	9550 } 10590 }	Infrared
Sodium	589	Yellow	Helium-neon	633	Red

PROPERTIES OF SELECTED MATERIALS

Substance	Density/ kg m^{-3}	Melting Point/ K	Boiling Point/ K	Specific Heat Capacity/ $\text{J kg}^{-1} \text{ K}^{-1}$	Specific Latent Heat of Fusion/ J kg^{-1}	Specific Latent Heat of Vaporisation/ J kg^{-1}
Aluminium	2.70×10^3	933	2623	9.02×10^2	3.95×10^5
Copper	8.96×10^3	1357	2853	3.86×10^2	2.05×10^5
Glass	2.60×10^3	1400	6.70×10^2
Ice	9.20×10^2	273	2.10×10^3	3.34×10^5
Glycerol	1.26×10^3	291	563	2.43×10^3	1.81×10^5	8.30×10^5
Methanol	7.91×10^2	175	338	2.52×10^3	9.9×10^4	1.12×10^6
Sea Water	1.02×10^3	264	377	3.93×10^3
Water	1.00×10^3	273	373	4.19×10^3	3.34×10^5	2.26×10^6
Air	1.29
Hydrogen	9.0×10^{-2}	14	20	1.43×10^4	4.50×10^5
Nitrogen	1.25	63	77	1.04×10^3	2.00×10^5
Oxygen	1.43	55	90	9.18×10^2	2.40×10^5

The gas densities refer to a temperature of 273 K and a pressure of $1.01 \times 10^5 \text{ Pa}$.

