

# X069/301

NATIONAL  
QUALIFICATIONS  
2002

WEDNESDAY, 22 MAY  
1.00 PM – 3.30 PM

PHYSICS  
HIGHER

### Read Carefully

- 1 All questions should be attempted.

### Section A (questions 1 to 20)

- 2 Check that the answer sheet is for Physics Higher (Section A).
- 3 Answer the questions numbered 1 to 20 on the answer sheet provided.
- 4 Fill in the details required on the answer sheet.
- 5 Rough working, if required, should be done only on this question paper, or on the first two pages of the answer book provided—**not** on the answer sheet.
- 6 For each of the questions 1 to 20 there is only **one** correct answer and each is worth 1 mark.
- 7 Instructions as to how to record your answers to questions 1–20 are given on page three.

### Section B (questions 21 to 30)

- 8 Answer questions numbered 21 to 30 in the answer book provided.
- 9 Fill in the details on the front of the answer book.
- 10 Enter the question number clearly in the margin of the answer book beside each of your answers to questions 21 to 30.
- 11 Care should be taken to give an appropriate number of significant figures in the final answers to calculations.



**DATA SHEET**  
COMMON PHYSICAL QUANTITIES

<i>Quantity</i>	<i>Symbol</i>	<i>Value</i>	<i>Quantity</i>	<i>Symbol</i>	<i>Value</i>
Speed of light in vacuum	$c$	$3.00 \times 10^8 \text{ m s}^{-1}$	Mass of electron	$m_e$	$9.11 \times 10^{-31} \text{ kg}$
Magnitude of the charge on an electron	$e$	$1.60 \times 10^{-19} \text{ C}$	Mass of neutron	$m_n$	$1.675 \times 10^{-27} \text{ kg}$
Gravitational acceleration	$g$	$9.8 \text{ m s}^{-2}$	Mass of proton	$m_p$	$1.673 \times 10^{-27} \text{ kg}$
Planck's constant	$h$	$6.63 \times 10^{-34} \text{ J s}$			

**REFRACTIVE INDICES**

The refractive indices refer to sodium light of wavelength 589 nm and to substances at a temperature of 273 K.

<i>Substance</i>	<i>Refractive index</i>	<i>Substance</i>	<i>Refractive index</i>
Diamond	2.42	Water	1.33
Crown glass	1.50	Air	1.00

**SPECTRAL LINES**

<i>Element</i>	<i>Wavelength/nm</i>	<i>Colour</i>	<i>Element</i>	<i>Wavelength/nm</i>	<i>Colour</i>
Hydrogen	656	Red	Cadmium	644	Red
	486	Blue-green		509	Green
	434	Blue-violet		480	Blue
	410	Violet	<i>Lasers</i>		
	397	Ultraviolet	<i>Element</i>	<i>Wavelength/nm</i>	<i>Colour</i>
	389	Ultraviolet	Carbon dioxide	9550 } 10590 }	Infrared
Sodium	589	Yellow	Helium-neon	633	Red

**PROPERTIES OF SELECTED MATERIALS**

<i>Substance</i>	<i>Density/ kg m<sup>-3</sup></i>	<i>Melting Point/ K</i>	<i>Boiling Point/ K</i>
Aluminium	$2.70 \times 10^3$	933	2623
Copper	$8.96 \times 10^3$	1357	2853
Ice	$9.20 \times 10^2$	273	...
Sea Water	$1.02 \times 10^3$	264	377
Water	$1.00 \times 10^3$	273	373
Air	1.29	...	...
Hydrogen	$9.0 \times 10^{-2}$	14	20

The gas densities refer to a temperature of 273 K and a pressure of  $1.01 \times 10^5 \text{ Pa}$ .

